Periodic table

Answers

KS3 science quick questions

This topic online

Activate 2: The Periodic Table; Activate 3: Atoms and the Periodic Table | Exploring Science 8F

Q	Answer	Remember
1	groups	Groups go up and down. Elements in a group behave in similar ways.
2	periods	Periods go across, like reading a line of text.
3	left	Metals fill the left and middle; non-metals sit up on the right.
4	metals	Roughly three quarters of all elements are metals.
5	alkali metals	They react with water to make alkaline solutions - hence the name.
6	halogens	Halogen comes from Greek words meaning "salt maker".
7	noble gases	Like nobles who keep to themselves, these gases rarely react with anything.
8	unreactive	Their unreactiveness makes them safe for lights, balloons and welding.
9	Mendeleev	His gaps were later filled by elements that matched his predictions.
10	chemical	Similar chemical properties is the whole reason elements are grouped together.
11	increases	Lithium fizzes gently in water, sodium whizzes about and potassium bursts into flame.
12	decreases	This is the opposite trend to Group 1 - fluorine is the most reactive halogen.
13	hydrogen	Alkali metal + water makes a metal hydroxide and hydrogen.
14	oil	Oil keeps air and moisture away from the metal.
15	chlorine	A tiny amount kills bacteria, which helps make tap water safe to drink.
16	increase	Gas, then liquid, then solid as you go down - a clear sign the boiling points rise.
17	more	Use the trend: the further down Group 1, the more violent the reaction.
18	bromine	A more reactive halogen displaces a less reactive one from its compound.
19	sodium	Period 3 starts with sodium, then magnesium, then aluminium.
20	He left gaps for elements not yet discovered, so that each element sat in a group with others	He trusted the patterns more than the list of known elements - and he was right.

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Q	Answer	Remember
	that had similar properties.	

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