Chemical energy

Answers

KS3 science quick questions

This topic online

Activate 1: Reactions; Activate 3: Chemical changes | Exploring Science 8E

Q	Answer	Remember
1	exothermic	Exo = exit: the energy leaves the reaction.
2	endothermic	Endo sounds like "enter": the energy goes in.
3	exothermic	All burning reactions give out heat, so they are exothermic.
4	increases	Energy is passed out to the surroundings, so they warm up.
5	endothermic	It takes heat from your skin, so it feels cold.
6	endothermic	Photosynthesis needs a constant supply of light energy.
7	catalyst	A catalyst can be used again and again.
8	no	You get the same mass of catalyst back at the end.
9	enzymes	Enzymes are biological catalysts - they work best at body temperature.
10	taken in	You must put energy in to pull atoms apart.
11	given out	Breaking takes in, making gives out.
12	thermometer	Read the temperature before and after, then find the difference.
13	8 °C	Change = end temperature - start temperature = 28 - 20.
14	endothermic	The temperature dropped, so the reaction took in heat.
15	physical change	Changes of state can be reversed easily and make no new substances.
16	less	The missing energy is the heat that was given out.
17	200 kJ	700 - 500 = 200. More given out than taken in means exothermic.
18	endothermic	More energy goes in than comes out, so the surroundings get colder.
19	activation energy	Striking the match gives that first push of energy; the burning then keeps itself going.
20	Catalysts make reactions go faster without being used up, so products are made more quickly and with less energy, which saves money.	Faster and cheaper, and the catalyst can be reused.

Free for classroom and home use: print, photocopy and share. Not for resale or re-hosting.