Metals and non-metals

Answers

KS3 science quick questions

This topic online

Activate 2: Metals and other materials | Exploring Science 8G, 9F

Q	Answer	Remember
1	good	Metals also conduct heat well, which is why pans are made of them.
2	shiny	Most metals are shiny when freshly cut, but some go dull as they react with air.
3	malleable	Malleable = can be bent or beaten into shape. Think of a "mallet".
4	ductile	Ductile is for wires; malleable is for sheets and shapes.
5	mercury	Mercury was once used in thermometers because it is a runny liquid metal.
6	poor	That is why plastic and rubber, made from non-metals, are used to cover wires.
7	carbon	Graphite is the odd one out - a non-metal that conducts.
8	hydrogen	Metal + acid = salt + hydrogen. The bubbles you see are hydrogen.
9	squeaky	The squeaky pop is hydrogen burning quickly with oxygen in the air.
10	metal oxide	Metal + oxygen = metal oxide, for example copper + oxygen = copper oxide.
11	rust	Rusting needs both water and oxygen - paint keeps them away.
12	reactivity series	Potassium is near the top; gold is right at the bottom.
13	potassium	Potassium reacts violently with water; iron only rusts slowly.
14	copper	Copper is below hydrogen in the reactivity series, so it does not react with dilute acids.
15	brittle	Solid non-metals tend to crumble or snap rather than bend.
16	copper	Iron is more reactive, so it pushes copper out of the solution.
17	no	Copper is less reactive than zinc, so it cannot take zinc's place.
18	zinc chloride	Hydrochloric acid always makes chlorides.
19	acidic	Non-metal oxides make acids in water; many metal oxides make alkalis.

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Q	Answer	Remember
20	Gold is so unreactive that it does not combine with other elements, but iron reacts with substances such as oxygen to form compounds.	The lower a metal is in the reactivity series, the more likely it is found on its own.

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